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6.02×10^{23}

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the amount of a

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substance. Avogadro's Number, the number of representative particles in a mole, 6.02×10^{23} .

Representative Particle.

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SI unit representing

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pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles. Quia - Chapter 10 "Chemical Quantities" Vocab

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Quantities. mole.

Avogadro's number.

representative

particle. molar mass.

the SI unit for

measuring the

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amount of a

substance. number of
representative

particles in a mole,

6.02×10^{23} . refers

to the species present
in a substance:

usually atoms, m....

the mass of one mole
of a pure substance.

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Chapter 10

Section 10.1 – The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro ' s number.
 $1 \text{ mole} = 6.02 \times 10^{23}$

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representative
particles

Quantities

Vocabulary
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Chemical Quantities

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Quantities. mole.

Avogadro's number.

representative

particle. molar mass.

the SI unit for

measuring the

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representative

particles in a mole,
 6.02×10^{23} . refers
to the species present
in a substance:

usually atoms, m....
the mass of one mole
of a pure substance.

chemical quantities

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10: Chemical

Quantities BASICS: •

The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he

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decided to use the

Quantities

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unit representing

6.02×10^{23}

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particles of a

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1 mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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